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Particle Mass number 19. Proton 20. Neutron 21. Electron 10 CHAPTER 3 TEST Relative charge Location MODERN CHEMISTRY HRW material copyrighted under notice appearing earlier in this work. Menu Print Name Date Class CHAPTER 3 TEST continued SHORT ANSWER Write the answers to the following questions in the space provided. 22.

Holt Modern Chemistry Chapter3 Practice Test

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Chapter 4 Modern Atomic Theory - An Introduction to Chemistry

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inorganic chemistry—the study of non-organic substances physical chemistry—the study of properties of matter, changes that occur in matter, and the relationships between matter and energy analytical chemistry—the identification of the composition of materials biochemistry—the study of the chemistry of living things

1 Matter and Change - HUBBARD'S CHEMISTRY

Heat is applied to the reactants. A chemical is dissolved in water. A chemical is in the liquid state. 2, Finish balancing the following equation: $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$. In each of the following formulas, write the total number of atoms present. 12 atoms 16 atoms 51 atoms 3×10^{24} atoms a. b. 8.02×10^{23} d. 6×10^{23} 4.

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I ordered two books for my high schooler's Chemistry class, "section reviews" and "study guide" by the same authors, to go with her text book "Modern Chemistry". I was very disapointed to find out that these two books are identical in the first 80 pages, the same title, the same contents, the same page number etc, but the cover is completely ...

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SHORT ANSWER Answer the questions in the space provided. 1. What is the difference between mass and weight? 2. Identify the elements and the number of atoms of each element in each of the following compounds:
BO₂ C₆H₁₂O₆ 3. How many pairs of electrons do the two oxygen atoms in an oxygen molecule share with each other? Explain your answer. 4.

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