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Unit 9: Thermochemistry Homework
Packet (70 points)

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A sample of copper absorbs 234 J of

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heat. If the temperature changed from 0.52°C to 24°C , what was the mass of the copper? 7. A metal was heated from 5.00°C to 15.0°C and was 145 g. If it absorbed 341 J of heat, what was the metal? $0.235 \text{ J/g}^{\circ}\text{C}$ SILVER 8.

Unit 9: Thermochemistry Homework Packet (70 points)

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This homework set is divided into lessons that go perfectly with my Powerpoint. It is in Microsoft Word, so you can easily make changes to suit your needs. Answer key is included.

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sound, and light. • Worksheet 1 C3.3
Heating Impacts: Heating a substance
increases its kinetic energy • Worksheet
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Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words ΔH is positive? I. $\text{H}_2\text{O}(\text{l}) + 10,5\text{kcal} \rightarrow \text{H}_2\text{O}(\text{g}) \Delta H1$ II. $2\text{NH}_3 + 22\text{kcal}$

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addition of 14.0 g of KOH to water:

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LO MAPPED - The FRQ's that are listed on a worksheet marked as 'LO MAPPED', are mapped to SPECIFIC LO's in the May 2019 CED. It should be noted that NONE of the questions prior to 2020 were actually mapped by the College Board to

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the May 2019 CED, since they were all written prior to that document being produced.

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This chapter introduces you to thermochemistry, a branch of chemistry that describes the energy changes that occur during chemical reactions. In some situations, the energy produced by chemical reactions is actually of greater interest to chemists than the material products of the reaction.

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5: Thermochemistry - Chemistry

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Thermochemistry deals with heat (energy) changes in chemical reactions. In chemical reactions heat is released or absorbed. If reaction absorbs heat then we call them endothermic reactions and if reaction release heat we call them exothermic reactions .

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Thermochemistry | Online Chemistry Tutorials

Experiment 10: Thermochemistry:
Hess's Law Purpose Determine the
enthalpy of the ionization reaction of
acetic acid: $\text{CH}_3\text{COOH}(\text{ac}) \rightarrow \text{CH}_3\text{COO}^- (\text{aq}) + \text{H}^+(\text{aq})$
Hess's Law will be used to calculate the
enthalpy change from two independent

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reactions. Background Information after the decimal point) is inserted through the hole on the cover.

Solved: Experiment 10: Thermochemistry: Hess's Law Purpose ...

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Match each item with the correct

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statement below. a. calorimeter d.
enthalpy b. calorie e. specific heat c.
joule f. heat capacity ____ 1. quantity of
heat needed to raise the temperature of
1 g of water by 1 C 2. ____ SI unit of
energy 3. ____ quantity of heat needed to
change the temperature of 1 g of a ...

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These high school chemistry worksheets are full of pictures, diagrams, and deeper questions covering all aspects of thermochemistry and states of matter! This unit is meant to cover the basics of enthalpy, specific heat problems, standard heat of formation, phase diagrams, time and temperature graphs, Hess' Law, and even a Born-Haber

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