

## Chemistry Solutions

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### Chemistry Solutions

Solution, in chemistry, a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is commonly applied to the liquid state of matter, but solutions of gases and solids are possible.

#### solution | Definition & Examples | Britannica

NCERT Solutions for Class 12 Chemistry: The NCERT solutions provided here will enhance the concepts of the students, as well as suggest alternative methods to solve particular problems to the teachers. The target is to direct individuals towards problem solving strategies, rather than solving problems in one prescribed format.

#### NCERT Solutions for Class 12 Chemistry (Updated for 2020-21)

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A solution consists of a solute and a solvent. The solute is the substance that is dissolved in the solvent. The amount of solute that can be dissolved in solvent is called its solubility. For example, in a saline solution, salt is the solute dissolved in water as the solvent.

#### Solution Definition in Chemistry - ThoughtCo

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#### Solutions - Chemistry LibreTexts

In chemistry, a solution is a special type of homogeneous mixture composed of two or more substances. In such a mixture, a solute is a substance dissolved in another substance, known as a solvent. The mixing process of a solution happens at a scale where the effects of chemical polarity are involved, resulting in interactions that are specific to solvation.

#### Solution - Wikipedia

Example of a Solution One example of a solution is salt water which is a mixture of water and salt. You cannot see the salt and the salt and water will stay a solution if left alone. Parts of a Solution. Solute - The solute is the substance that is being dissolved by another substance. In the example above, the salt is the solute.

#### Chemistry for Kids: Solutions and Dissolving

NCERT Solutions for Class 12 Chemistry (all chapters) are provided chapter-wise on this page. Students can access solutions for each individual Class 12 chemistry chapter by following the links tabulated below. The NCERT solutions provided here are free for all users to download as a PDF or to view online.

#### NCERT Solutions for Class 12 Chemistry (All Chapters) with PDF

NCERT TEXTBOOK QUESTIONS SOLVED. 2.1. Calculate the mass percentage of benzene (C<sub>6</sub>H<sub>6</sub>) and carbon tetrachloride (CCl<sub>4</sub>) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride. Ans: Mass of solution = Mass of C<sub>6</sub>H<sub>6</sub> + Mass of CCl<sub>4</sub> = 22 g + 122 g = 144 g Mass % of benzene =  $\frac{22}{144} \times 100 = 15.28\%$

#### NCERT Solutions For Class 12 Chemistry Chapter 2 Solutions

Such solutions are called electrolytes. If the dissociation of ions is complete, the solution is a strong electrolyte. If the dissociation is only partial, the solution is a weak electrolyte. Solutions of molecules do not conduct electricity and are called nonelectrolytes. Solutions have properties that differ from those of the pure solvent.

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Chemistry of solutions 1. D' CHEMISTRY OF SOLUTIONS 2. SOME DEFINITIONS A solution is a mixture of 2 or more substances in a single phase. One constituent is usually regarded as the SOLVENT and the others as SOLUTES

#### Chemistry of solutions - SlideShare

where P solution is the observed vapor pressure of the solution, X solvent is the mole fraction of solvent, and P<sub>o</sub> solvent is the vapor pressure of the pure solvent. Any solution that obeys Raoult's law is called an ideal solution. Raoult's law is to solutions what the ideal gas law is to gases.

#### Solution Chemistry

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solution, in chemistry, homogeneous mixture mixture, in chemistry, a physical combination of two or more pure substances (i.e., elements or compounds). A mixture is distinguished from a compound, which is formed by the chemical combination of two or more pure substances in a fixed, definite proportion.

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